

Pyrazolate-Based Dinuclear α -Diimine-Type Palladium(II) and Nickel(II) Complexes – a Bimetallic Approach in Olefin Polymerisation

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Abstract: A series of compartmental pyrazole/imine dinucleating ligands **HL**¹–**HL**⁴ have been prepared with different backbone substituents and different steric bulk of the appending arylimine groups, and two of them have been structurally characterised. With PdCl₂ all ligands form the anticipated pyrazolate-bridged bimetallic complexes LPd₂Cl₃ (**1a**–**4a**) that feature square-planar metal ions, as is confirmed by the X-ray crystallographic analyses of L¹Pd₂Cl₃ (**1a**) and L³Pd₂Cl₃ (**3a**). With NiBr₂(dme), however, oligonuclear systems [LNi₂Br₃]_x (**1b**–**4b**) are formed according to mass spectrometric findings, which is corroborated by the crystallographic structure of [L³Ni₂Br₃]₃. Complexes **1a**–**4a** can be described as bimetallic versions of Brookhart-type α -diimine palladium complexes, where dissociation into mononuclear species is prevented by the dinucleating scaffold and

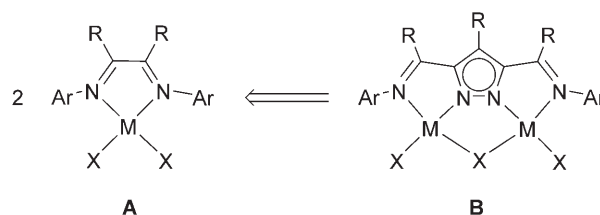
the proximate metal ions are suitably positioned to work in concert during substrate transformation. Upon activation of the complexes with MAO and exposure to ethylene, polyethylene is formed. Whereas the palladium complexes display moderate activities, nickel complexes are very active. From structure/activity correlations it is evident that the presence of backbone substituents at the pyrazolate scaffold as well as bulky *ortho* aryl substituents is advantageous for polymerisation. Overall, activities of the Ni complexes and the microstructure of the polymer obtained (total branching, *T_m* and molecular weights) are still rather similar to the data reported previously for mononuclear cationic diimine nickel complexes.

Keywords: dinuclear complexes; N ligands; nickel; olefin polymerisation; palladium

Introduction

Since Brookhart et al. introduced nickel and palladium complexes of α -diimine ligands^[1] with bulky aryl substituents as efficient pre-catalysts for olefin polymerisation,^[2] these systems and various derivatives thereof have been studied extensively.^[3] Most of the catalytic chemistry of such type **A** compounds is based on the presence of two active coordination sites at the single metal centre, which prompted us to devise novel bimetallic type **B** complexes where two adjacent metal ions might work in concert (Scheme 1). The basic concept, i.e., the use of a pyrazolate bridge to (in a formal sense) link two N-based ligand compartments in a preorganised dinucleating scaffold, has been outlined previously

and has proven successful for the investigation of bimetallic versions of N-donor-functionalised cyclopentadienyl compounds^[4] as well as bimetallic versions of classical Werner-type coordination compounds.^[5,6]



Scheme 1.

A synthetic route to a multidentate pyrazolate ligand **HL**³ that bears bulky 2,6-diisopropylphenyl substituents at the imine N atoms and that appears suitable for the preparation of type **B** systems has been described recently,^[7,8] and a first dinickel(II) complex of this ligand has been communicated.^[8] The latter was found to form an unusual hexanuclear aggregate $[\text{L}^3\text{Ni}_2\text{Cl}_3]_3$ in the solid state that incorporates a μ_6 -chloride ion in the centre of a trigonal prism of nickel(II) ions. Here we report (i) the synthesis of a series of dinucleating pyrazole/imine ligands with either H atoms ($\text{R}=\text{H}$, such as in **HL**³) or more bulky methyl groups ($\text{R}=\text{Me}$) attached to the backbone, and with different steric bulk of the appended aryl groups, (ii) the molecular structures of nickel(II) and palladium(II) complexes of these ligands, and (iii) the results of ethylene polymerisation studies using the new complexes and MAO as a cocatalyst. Potential cooperative effects in such highly preorganised bimetallic complexes may offer new options in metal-mediated catalytic reactions.^[9] In particular, few bimetallic catalyst systems have hitherto been used in olefin polymerisation,^[10–12] and most of them provide a flexible linker between two Cp-derived entities.^[10] In contrast, the present type **B** systems offer a highly preorganised dinuclear array in which each subunit closely resembles that of the well-studied Brookhart-type α -diimine systems **A**. Because of the compartmental design of the pyrazolate-based ligand scaffold, the dinuclear arrangement is expected to stay intact under typical conditions for olefin polymerisation (i.e., substitution or abstraction of coligands X upon reaction with MAO cocatalyst). In addition, all accessible coordination sites are directed towards the same side of the bimetallic set-up and are thus favourably positioned for cooperative action.

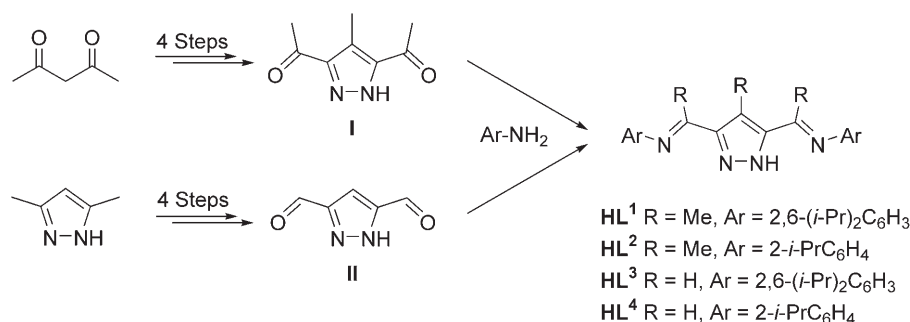
Results and Discussion

Synthesis and Structural Characterisation of Ligands

For the preparation of type **I** pyrazole-diimine ligands with either H atoms or methyl groups in the backbone,

efficient syntheses based on cheap and commercially available starting materials had to be developed. The pyrazole derivatives **I** and **II** with carbonyl substituents in the 3- and 5-positions of the heterocycle were prepared with slight modifications of the procedures reported in the literature^[13,14] and were used in the subsequent Schiff base formation (Scheme 2). Proper conditions turned out to be critical in order to achieve reasonable yields in the final condensation step and to assure formation of the difunctionalised products. Preparation of **HL**¹–**HL**⁴ was carried out by the condensation of 6.0 equivalents of the corresponding substituted aniline with 1.0 equivalent of diketone **I** or dialdehyde **II** in the presence of a catalytic amount (10 mol %) of *p*-toluenesulfonic acid in toluene solution, using a Dean–Stark apparatus for the removal of water. After reacting for 72 h, compounds **HL**¹–**HL**⁴ are isolated as colourless solids by column chromatography in yields up to 85%.

Crystals of **HL**¹ suitable for X-ray structural determination were grown from an ethanol solution layered with *n*-hexane, while crystals of **HL**⁴ were obtained from CHCl_3 . The molecular structures are shown in Figures 1 and 2. In both cases, iminomethyl groups attached to the 3- and 5-positions of the heterocycle are roughly coplanar with the pyrazole ring, and both molecules adopt an *anti*-conformation with respect to the orientation of the two side arms attached to the pyrazole: while N(4) is located close to the pyrazole-N(2), N(3) is directed towards the backside away from pyrazole-N(1). Hydrogen bonding between pairs of nitrogen atoms results in the formation of dimeric structures in the solid state, but H-bonding patterns are different for the two compounds. In **HL**¹ only the nitrogen atoms N(1) and N(2) of the pyrazole are involved in dimer formation, and the resulting dimer has crystallographically imposed inversion symmetry. In contrast, two crystallographically independent molecules constitute the dimeric structure of **HL**⁴, in which hydrogen bonding occurs between pairs of pyrazole [N(2), N(12)] and imine [N(4), N(14)] nitrogen atoms. While the dimer motif observed for **HL**¹ is quite common for solid-state structures of pyrazole derivatives,^[15] the arrangement seen for **HL**⁴ that involves other H-bond acceptors in the same molecule is rare.



Scheme 2. Ligands used in this study.

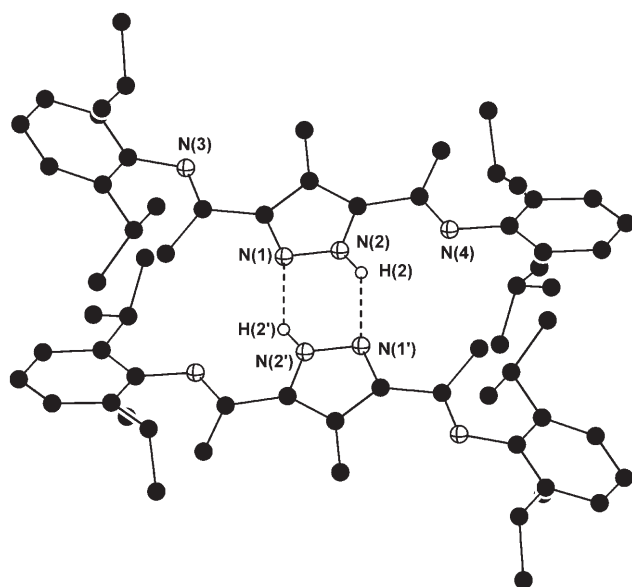


Figure 1. Hydrogen bond structure of **HL**¹. In the interest of clarity most hydrogen atoms have been omitted. Selected atom distances (Å) and angles (°): N(2)⋯N(1') 2.898(3); N(2)–H(2)⋯N(1') 144(2). Symmetry operation used to generate equivalent positions: (') 1-x, 1-y, 1-z.

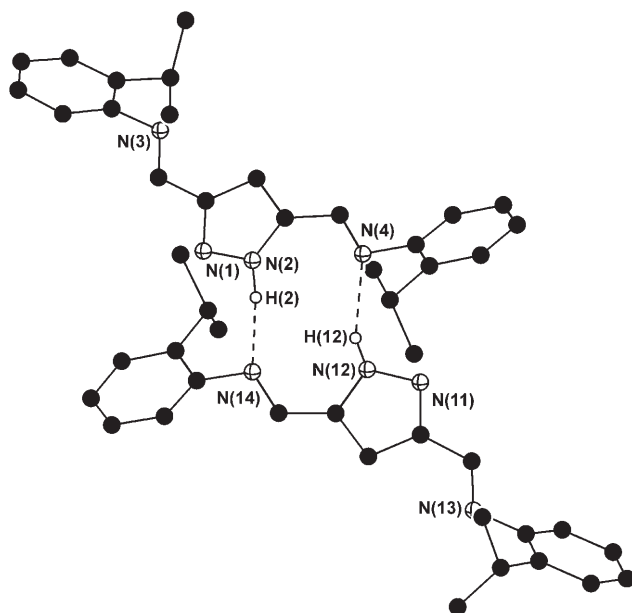
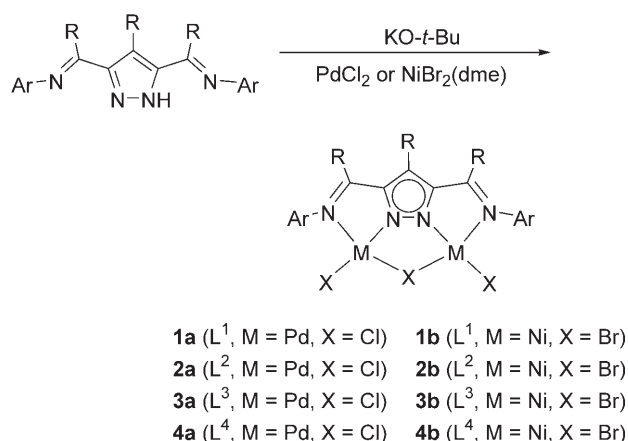


Figure 2. Hydrogen bond structure of **HL**⁴. In the interest of clarity most hydrogen atoms have been omitted. Selected atom distances (Å) and angles (°): N(2)⋯N(14) 2.940(9), N(12)⋯N(4) 2.947(9); N(2)–H(2)⋯N(14) 168(5), N(12)–H(12)⋯N(4) 155.

Synthesis and Structural Characterisation of Complexes

For the synthesis of dipalladium(II) and dinickel(II) complexes, the respective ligand was first deprotonated



Scheme 3. Syntheses of the complexes.

by means of a stoichiometric amount of KO-*t*-Bu in THF to generate the pyrazolate and then treated with 2 equivalents of either PdCl₂ or NiBr₂(dme) (Scheme 3). After stirring at room temperature for 72 h, the products were isolated as air-stable orange-red (Pd) or yellow (Ni) solids in moderate to good yield. They have been characterised by elemental analysis, mass spectrometry and IR spectroscopy, as well as by NMR spectroscopy for **1a** and **3a**. Complexes of [L^2][−] and [L^4][−] are poorly soluble in most common solvents, and all nickel complexes turned out to be paramagnetic in solution.

Upon complexation, the characteristic C=N stretching frequency of the appended imine groups shifts from 1625–1642 cm^{−1} in the free ligands to 1530–1561 cm^{−1} (**1a–4a**) or 1555–1610 cm^{−1} (**1b–4b**) in the complexes. For some of the nickel complexes, several strong bands in the latter range are observed, suggesting the presence of different isomers or species of different nuclearity. ESI mass spectra of acetonitrile solutions of the palladium complexes show prominent signals for the expected bimetallic compounds devoid of one of the Cl ligands, i.e., for the ions [LPd_2Cl_2]⁺. In the case of **1a** and **2a**, however, peaks are also discernible for, *inter alia*, [$L_2Pd_4Cl_5$]⁺, again indicating the presence of some aggregation equilibria in solution that give rise to species of higher nuclearity. This is even more pronounced for the corresponding nickel complexes, where peaks not only for [$L_2Ni_4Br_5$]⁺, but also for hexametallic species [$L_3Ni_6Br_x$]⁺ are detected in some cases. The positive ion FAB mass spectrum of **3b** is shown in Figure 3 as an example. While the most intense signal at m/z = 1517 corresponds to the ion [$L_2Ni_4Br_5$]⁺, a prominent signal for [$L_3Ni_6Br_8$]⁺ with the expected isotopic distribution pattern around m/z = 2316 suggests the existence of trimeric assemblies [$L^3Ni_2Br_3$]₃ as a major species. This is confirmed by X-ray crystallography (see below).

Single crystals of **1a** and **3a** were grown from a dichloromethane solution layered with light petroleum. Molecular structures are shown in Figures 4 and 5, respectively. As anticipated, the compartmental ligand

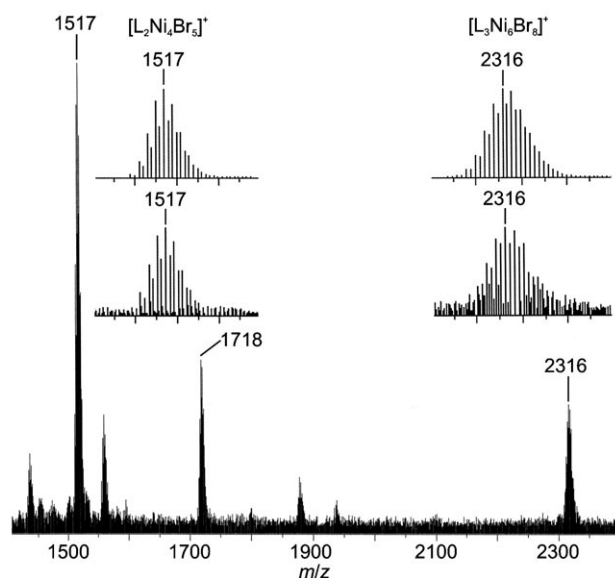
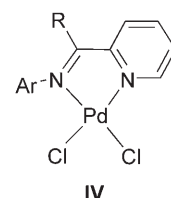
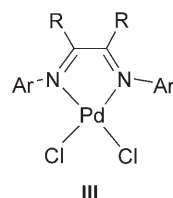


Figure 3. Part of the FAB mass spectrum of **3b**; the inset shows the experimental (lower) and theoretical (upper) isotopic distribution for tetranuclear $[L_2Ni_6Br_8]^+$ and hexanuclear $[L_3Ni_6Br_8]^+$.

scaffolds form highly preorganised dinuclear palladium(II) complexes in which each metal is bound to one of the pyrazolate-N and to the adjacent imine-N. A bridging chloride as well as a terminal chloride at each palladium complete the bimetallic array. All metals ions are found in square planar environment [distance to the least squares planes of the four surrounding ligand atoms for Pd(1), **1a**: 0.008(1) Å, **3a**: 0.012(1) Å; for Pd(2) **1a**: 0.011(1) Å, **3a**: 0.039(1) Å] with their coordination planes lying roughly within the plane of the pyrazolate heterocycle. The appended bulky aryl groups are oriented perpendicular to this plane, and thus the isopropyl substituents block much of the outer space above and below the axial metal coordination sites. The only difference between **1a** and **3a** is the presence of the backbone methyl groups in the former case ($R = \text{Me}$ versus $R = \text{H}$). Variations of the backbone substituents at the imine-C have been shown to greatly affect olefin polymerisation by mononuclear type **A** α -diimine catalysts.^[16,17] In the solid state structures investigated here, the presence of methyl groups ($R = \text{Me}$) in **1a** has no significant influence on the bond distances and angles compared to **3a**. In both compounds the Pd–Cl distances to the bridging chloride are about 0.1 Å longer than the distances to the terminal ones. A difference of the same magnitude can be found for the metal nitrogen bonds, where the distances to the imine-N are about 0.1 Å longer than those to the pyrazolate-N. The former values are quite similar to the Pd–N_{imine} distances found in mononuclear diazabutadiene complexes of PdCl_2 (**III**)^[18] or some related pyridylimine complexes **IV**.^[19,20] In contrast, the bonds between Pd and the anionic pyrazolate in **1a** and **3a** [1.909(5)–1.931(5) Å] are significantly shorter than

the Pd–N_{pyridine} bonds in type **IV** systems (in the range 2.02–2.12 Å). However, Pd–Cl distances *trans* to the pyrazolate-N or pyridine-N are all found in the narrow range 2.26–2.30 Å, suggesting a similar *trans*-influence of the two heterocyclic moieties. The Pd···Pd distances in **1a** (3.76 Å) and **3a** (3.74 Å) are longer than in a dinuclear diazabutadiene palladium complex with two μ -chloro bridges $\{[(\text{RNC}(\text{Me})\text{C}(\text{Me})\text{NR})\text{Pd}(\mu\text{-Cl})_2]_2\}^{2+}$, $R = 2,6\text{-}(i\text{-Pr})_2\text{C}_6\text{H}_3$, $d(\text{Pd}\cdots\text{Pd}) = 3.41$ Å,^[21] but are typical for pyrazolate-based bimetallic systems.^[22]



The complex with NiBr_2 and $[L^3]^-$, **3b** (Figure 6), differs substantially from the analogous square planar Pd complexes **1a** or **3a**. In **3b**, three bimetallic $\{L^3Ni_2Br_3\}$ building blocks have assembled to form a hexanuclear array, in accordance with the mass spectrometric findings discussed above. All metal ions are found in distorted octahedral coordination spheres. The cluster consists of two almost face-to-face arranged $Ni_3(\mu\text{-Br})_3$ rings (distance of the two ring centroids 4.11 Å, rotation angle of the two rings 8°) that are capped by two $\mu_6\text{-Br}$ atoms and are held together by three pyrazolate ligands acting as outside clips. A unique $\mu_6\text{-Br}$ atom is nested inside the

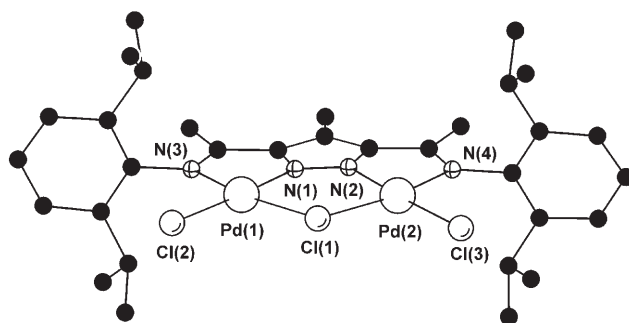


Figure 4. Molecular structure of **1a**. In the interest of clarity all hydrogen atoms have been omitted. Selected atom distances (Å) and angles (°): Pd(1)–N(1) 1.909(5), Pd(1)–N(3) 2.050(6), Pd(1)–Cl(1) 2.396(2), Pd(1)–Cl(2) 2.276(2), Pd(2)–N(2) 1.930(6), Pd(2)–N(4) 2.051(6), Pd(2)–Cl(1) 2.372(2), Pd(2)–Cl(3) 2.289(2), Pd(1)···Pd(2) 3.764(1); N(1)–Pd(1)–N(3) 78.2(2); N(1)–Pd(1)–Cl(1) 87.6(2), N(1)–Pd(1)–Cl(2) 175.2(2), N(3)–Pd(1)–Cl(2) 97.0(2), Cl(1)–Pd(1)–N(3) 165.8(2), Cl(1)–Pd(1)–Cl(2) 97.17(7), Pd(1)–Cl(1)–Pd(2) 104.23(7), N(2)–Pd(2)–N(4) 77.6(2), N(2)–Pd(2)–Cl(1) 88.1(2), N(2)–Pd(2)–Cl(3) 175.9(2), N(4)–Pd(2)–Cl(1) 165.6(2), N(4)–Pd(2)–Cl(3) 98.4(2), Cl(1)–Pd(2)–Cl(3) 96.01(7).

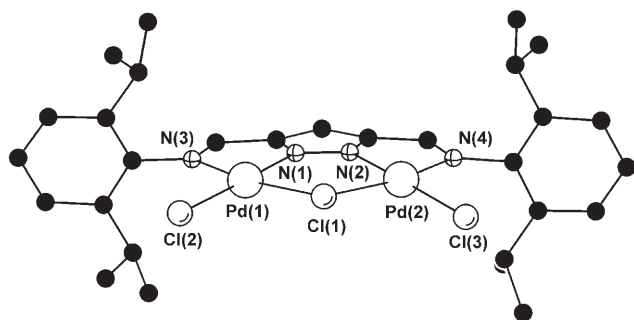


Figure 5. Molecular structure of **3a**. In the interest of clarity all hydrogen atoms have been omitted. Selected atom distances (Å) and angles (°): Pd(1)–N(1) 1.931(5), Pd(1)–N(3) 2.046(5), Pd(1)–Cl(1) 2.385(2), Pd(1)–Cl(2) 2.291(2), Pd(2)–N(2) 1.926(5), Pd(2)–N(4) 2.048(5), Pd(2)–Cl(1) 2.382(2), Pd(2)–Cl(3) 2.271(2), Pd(1)···Pd(2) 3.740(1); N(1)–Pd(1)–N(3) 77.4(2), N(1)–Pd(1)–Cl(1) 88.8(2), N(1)–Pd(1)–Cl(2) 174.9(2), N(3)–Pd(1)–Cl(2) 97.7(2), Cl(1)–Pd(1)–N(3) 166.2(2), Cl(1)–Pd(1)–Cl(2) 96.10(6), Pd(1)–Cl(1)–Pd(2) 103.37(6), N(2)–Pd(2)–N(4) 77.7(2), N(2)–Pd(2)–Cl(1) 88.7(2), N(2)–Pd(2)–Cl(3) 174.3(2), N(4)–Pd(2)–Cl(1) 166.3(2), N(4)–Pd(2)–Cl(3) 98.4(2), Cl(1)–Pd(2)–Cl(3) 95.24(6).

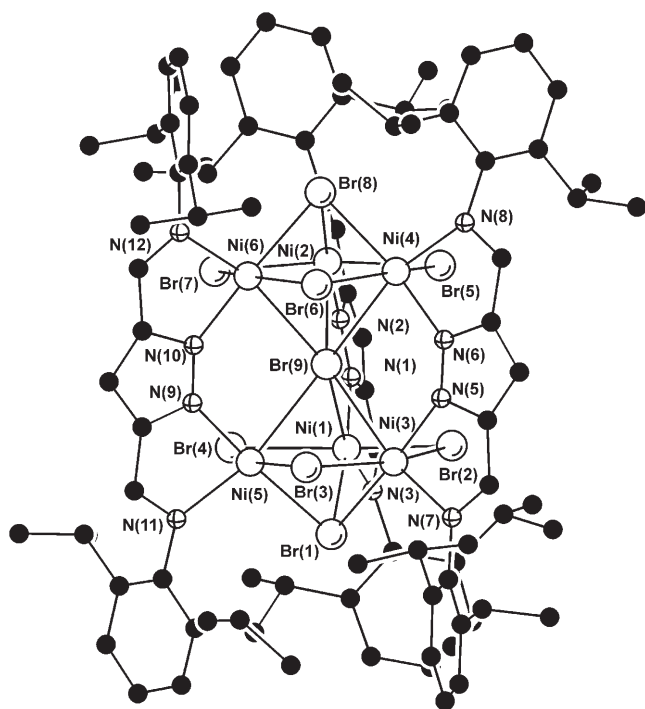


Figure 6. Molecular structure of **3b**. In the interest of clarity all hydrogen atoms have been omitted. Selected atom distances (Å): Ni–(μ -Br) 2.506(1)–2.782(1), Ni–(μ_3 -Br) 2.540(1)–2.593(1), Ni–Br(9) 2.782(1)–2.864(1), Ni–N_{pyrazolate} 1.979(5)–1.992(5), Ni–N_{imine} 2.098(5)–2.122(5), Ni···Ni within the Ni₃Br₃ rings: 3.265(1)–3.309(1), Ni···Ni bridged by pyrazolate: 4.191(1)–4.238(1).

cage-like structure. The complex is iso-structural with the recently communicated structure of $[\text{L}^3\text{Ni}_2\text{Cl}_3]_3$ except for the different halide.^[8] It should be noted that μ_6 -halide ligands bound to six metal atoms are very rare in discrete molecules, and for bromide only few structures of Hg compounds featuring a μ_6 -Br atom have been described in the literature to date.^[23]

Ethylene Polymerisation

Palladium(II) and nickel(II) complexes **1a–4a** and **1b–4b** were studied as catalyst precursors for ethylene polymerisation, with MAO as a cocatalyst. Results are summarised in Table 1. In all experiments the precatalyst complexes were dissolved in toluene and activated by MAO. Upon addition of MAO the colour of the solution changes from orange to dark red in the case of palladium(II) complexes **1a–4a** and from yellow to red for nickel(II) complexes **1b–4b**. The red colour for the latter systems is indicative of the formation of square-planar low-spin nickel(II) species upon methylation and suggests that the hexanuclear aggregates observed in the solid state break down under these conditions, presumably giving rise to discrete type **B** dinuclear species.

Initial polymerisation runs were carried out without temperature control in order to obtain a rough estimate about the range of activities covered by the novel catalyst precursors. For all ligand systems the nickel complexes are significantly more active than the corresponding palladium complexes, and for both series of complexes the order of activity is **1** > **2** > **3** > **4**. Aryl groups with more sterically demanding substituents [2,6-(*i*-Pr)₂C₆H₃ in **1** and **3** versus 2-*i*-PrC₆H₄ in **2** and **4**] and methyl groups at the diimine backbone (in **1** and **2**) are obviously advantageous, as was previously reported for mononuclear diimine complexes.^[2a,24]

The temperature inside the reactor was continuously monitored during these screening experiments, and the temperature pattern over the first 60 min is shown in Figure 7 for selected catalyst systems. It is apparent that polymerisation by the nickel catalysts generally is much more exothermic than for the corresponding palladium species. In particular, ethylene polymerisation by **1b**/MAO is very exothermic under the reaction conditions studied. For all systems investigated, however, a maximum in the temperature curve is observed, suggesting a decreasing polymerisation rate due to thermal decomposition of the catalysts at higher temperatures. The catalyst decay mechanism is unknown at present, but temperature sensitivity and decomposition above 60 °C have also been described as a general limitation for mononuclear type **A** Pd(II)- α -diimine or Ni(II)- α -diimine catalysts.^[24] In addition, limiting monomer diffusion into more viscous polymer solutions might also play a role.

For subsequent experiments with the most active palladium complex (**1a**) as well as with all nickel complexes,

Table 1. Ethylene polymerisation results.^[a]

Entry	Precatalyst complex	MAO/cat.	Cat. amount [μmol]	Temp. control	Yield PE [g]	Activity [kg _{PE} ·mol _{cat} ⁻¹ ·h ⁻¹ ·bar ⁻¹]
1	1a	1000	5.0	No	0.59	47
2	1a	1000	5.0	18 °C	1.83	146
3	2a	1000	5.0	No	0.47	37
4	3a	1000	5.0	No	0.42	33
5	4a	1000	5.0	No	0.24	19
6	1b	1000	5.0	No	5.43	434
7	1b	500	2.5	18 °C	2.02	323
8	1b	1000	2.5	18 °C	3.48	557
9	1b	2000	2.5	18 °C	3.54	566
10	1b	3000	2.5	18 °C	3.87	619
11	2b	1000	5.0	No	3.11	249
12	2b	1000	2.5	18 °C	2.51	402
13	3b	1000	5.0	No	1.40	112
14	3b	1000	2.5	18 °C	1.47	235
15	4b	1000	5.0	No	0.89	71
16	4b	1000	2.5	18 °C	1.18	189

^[a] Experimental conditions: 200-mL stirred glass reactor, 80 ml toluene, $p(\text{C}_2\text{H}_4)=2.5$ bar, reaction time 1 h.

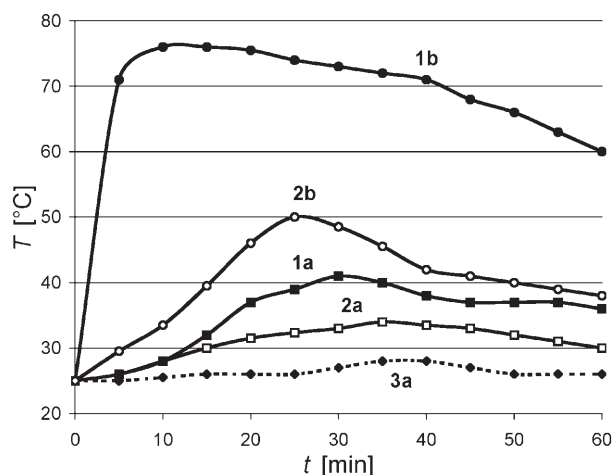


Figure 7. Temperature variation during ethylene polymerisation in the absence of temperature control [conditions: $p(\text{C}_2\text{H}_4)=2.5$ bar, MAO/precursor complex **1** to **4**=1000, 80 mL toluene, 5 μmol catalyst precursor].

the temperature inside the reactor was controlled by an external thermostat bath and was held roughly constant at 18 °C (Table 1). Overall structure/activity correlations remain valid under these conditions, i.e., methyl substituents at the ligand backbone as well as bulky groups at the aryl rings increase the polymerisation rate. Variation of the cocatalyst/catalyst ratio (MAO/cat) does have an effect on activity, although this effect levels off at ratios > 1000 (entries 7–10, Table 1). In order to evaluate the catalyst stability over time, in a series of experiments with the most active complex **1b** the reaction mixtures were quenched after different time intervals and polymer yields were determined. Figure 8 reveals that activity remains high and fairly constant

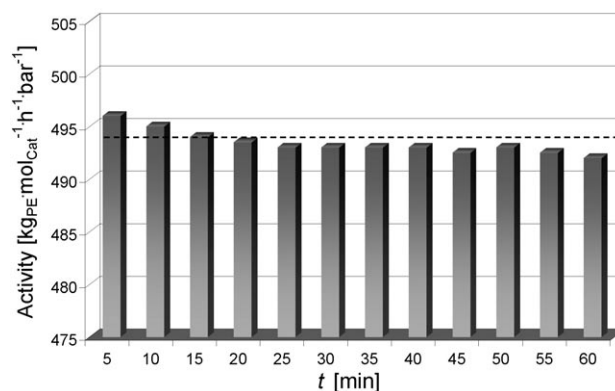


Figure 8. Long-time polymerisation activity for $[\text{L}^1\text{Ni}_2\text{Br}_3]_x$ (**1b**) [conditions: $p(\text{C}_2\text{H}_4)=2.5$ bar, 5 μmol **1b**, MAO/cat.=1000, 80 mL toluene, temperature controlled at 18 °C].

over at least 1 h, confirming good stability of the active species under thermostat conditions. The observed activities in this series of experiments also confirm a good reproducibility of the catalytic polymerisations.

Comparison with literature data reveals that ethylene polymerisation activity of the most active system **1b**/MAO is of the same order of magnitude as for the related mononuclear catalyst **V** [which was reported to produce 420 kg_{PE}/(mol_{Ni}·h) at 35 °C and 1 bar ethylene].^[24]

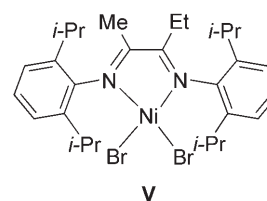


Table 2. Polymer properties.

Precatalyst complex	Entry in Table 1	M_n [g/mol]	Polydispersity M_w/M_n	T_m [°C]	Branches per 1000 C atoms ^[a]
1b	8	3.3×10^5	2.4	74	56
3b	14	3.7×10^5	2.7	76	51
V ^[b]	—	1.7×10^5	1.6	−2	101
V ^[c]	—	7.7×10^5	1.7	96	28

^[a] Determined by ¹H NMR spectroscopy.

^[b, c] Results from ref.^[24] Polymers prepared at 35 °C, MAO/cat = 1000 to 3000, 100 to 200 mL toluene, 1 to 2 μ mol Ni, ethylene pressure for [b]: 1 atm, [c]: 15 atm.

Polymer properties were studied for two selected samples, prepared with nickel(II) catalysts **1b** and **3b** that carry the same bulky 2,6-diisopropylphenyl imine substituents but differ in the backbone substitution. High temperature GPC shows the polymers to be of high molecular weight (Table 2). Number average molecular weights of several 10^5 g·mol^{−1} were determined *versus* linear polyethylene standards. Molecular weight distributions and the monomodality of the GPC traces indicate a well-behaved single site-type nature of the active species. A total branching of 50 to 60 branches per 1000 C atoms results in a much decreased melting point and crystallinity compared to linear polyethylene. The backbone substitution pattern (Me or H) of the catalyst appears to have no strong impact on the polymer microstructure. Overall, the polymer properties are rather similar to those obtained under similar conditions with the mononuclear diimine complex **V** (Table 2, properties of polymers prepared with **V** at 1 atm and at 15 atm ethylene pressure are given for comparison with the polymers prepared in this work at 2.5 atm). Note that branching decreases and correspondingly the melt temperature increases at higher ethylene pressure during polymerisation.

Conclusions

A set of pyrazolate-based ligands with appended arylimino side arms has been prepared that provide two adjacent bidentate metal ion binding compartments similar to the classical α -diimine type. In the case of bulky *ortho* aryl groups, these ligands form highly preorganised dipalladium(II) complexes with square-planar metal ions. Nickel (II) complexes tend to form oligomeric species both in solution and in the solid state. Upon activation of the new complexes with MAO and exposure to ethylene, polyethylene is formed. The following conclusions regarding ethylene polymerisation can be drawn from this work:

(i) The dinuclear complexes can be viewed as bimetallic versions of Brookhart-type α -diimine olefin polymerisation precatalysts, where the two adjacent metal ions are suitably positioned for cooperative action.

(ii) Dinuclear nickel complexes show high activity and are significantly more active than the corresponding palladium systems, in close analogy to observations for the mononuclear catalysts.

(iii) Increased steric bulk of the aryl *ortho* substituents and the presence of substituents at the ligand backbone lead to higher activity.

(iv) Polymer microstructures are rather similar to those obtained with related mononuclear α -diimine complexes under similar conditions, even though dissociation into monomeric complexes is prevented by the dinucleating ligand scaffold for the new bimetallic pyrazolate/imine systems. With bulky 2,6-diisopropylphenyl-substituted systems, M_n values are in the order of 3.5×10^5 g/mol, but no effect of different backbone substituents (H *versus* Me) is discernible. Quite extensive branching leads to polymers with relatively low crystallinity, while molecular weight distributions indicate a well-behaved single-site type nature of the active species.

Future work will focus on beneficial modifications of the ligand scaffold, on polymerisation of monomers other than ethylene, and on mechanistic aspects of the polymerisation process. It remains to be elucidated whether (and how) the proximate metal ions in the new complexes work in concert during substrate turnover.

Experimental Section

General

Compounds **I**,^[13] **II**^[14] and **HL**^[37] were synthesised according to published procedures. In order to improve the conversion yield in the last step of the synthesis of **II**, a mixture of water and ethanol (1:4) instead of pure water was used as the solvent. Characterisation data for **HL**¹, **HL**², **HL**⁴, **1a–4a**, and **1b–4b** can be found in the Supporting Information. Solvents were dried under reflux (CaH₂ for dichloromethane; Na/K-alloy for toluene, diethyl ether, THF and light petroleum) and distilled prior to use. All reactions were carried out under nitrogen atmosphere using standard Schlenk techniques. Anilines were purchased from Aldrich and purified by distillation prior to use. PdCl₂ and NiBr₂(dme) were purchased from Aldrich and used as received. Ethylene (99.9%) was purchased from Messer

Group GmbH. Methyl aluminoxane (MAO) was obtained from Albemarle® Europe SPRL as a 30 wt % solution in toluene.

3,5-Bis(2,6-diisopropylphenyliminoacetyl)-1*H*-4-methylpyrazole (**HL**¹)

A solution of diketone **I** (1.66 g, 10 mmol) and 2,6-diisopropylaniline (10.62 g, 60 mmol) in toluene (150 mL) with a catalytic amount of *p*-toluenesulfonic acid (0.20 g) was heated to reflux for 72 h using a Dean–Stark apparatus for the removal of water. After the solvent had been removed under reduced pressure, the remaining white powder was purified by column chromatography (light petroleum:Et₂O = 9:1), dried under vacuum and finally crystallised from ethanol. Single crystals were obtained from an ethanol solution layered with *n*-hexane. Yield: 3.70 g (76%).

3,5-Bis(2-isopropylphenyliminoacetyl)-1*H*-4-methylpyrazole (**HL**²)

A solution of diketone **I** (1.66 g, 10 mmol) and 2-isopropylaniline (6.42 g, 60 mmol) in toluene (150 mL) with a catalytic amount of *p*-toluenesulfonic acid (0.20 g) was heated to reflux using a Dean–Stark apparatus for the removal of water for 72 h. After the solvent had been removed under reduced pressure, the remaining white powder was purified by column chromatography (light petroleum:Et₂O = 9:1) and dried under vacuum. Yield: 3.27 g (82%).

3,5-Bis(2-isopropylphenyliminomethyl)-1*H*-pyrazole (**HL**⁴)

A solution of dialdehyde **II** (1.24 g, 10 mmol) and 2-isopropylaniline (6.42 g, 60 mmol) in toluene (150 mL) with a catalytic amount of *p*-toluenesulfonic acid (0.20 g) was heated to reflux using a Dean–Stark apparatus for the removal of water for 72 h. After the solvent had been removed under reduced pressure, the remaining white powder was purified by column chromatography (light petroleum:Et₂O = 9:1), dried under vacuum and finally crystallised from ethanol. Single crystals were obtained from a CHCl₃ solution. Yield: 3.29 g (92%).

[**L**¹Pd₂Cl₃] (**1a**)

To a stirred solution of **HL**¹ (242 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), PdCl₂ (177.4 mg, 1.0 mmol) was added. The suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting orange powder was taken up in CH₂Cl₂ (40 mL), filtered and the solution layered with light petroleum to gradually afford single crystals of **1a**. Yield: 324 mg (80%).

[**L**²Pd₂Cl₃] (**2a**)

To a stirred solution of **HL**² (200 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), PdCl₂ (177.4 mg, 1.0 mmol) was added. The suspension was stirred for 72 h at

room temperature and then evaporated to dryness. The resulting red powder was taken up in CH₂Cl₂ (40 mL) and filtered. All volatile material was then evaporated under reduced pressure to give **2a**. Yield: 280 mg (78%).

[**L**³Pd₂Cl₃] (**3a**)

To a stirred solution of **HL**³ (221 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), PdCl₂ (177.4 mg, 1.0 mmol) was added. This suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting orange powder was taken up in CH₂Cl₂ (40 mL), filtered and the solution layered with light petroleum to gradually afford single crystals of **3a**. Yield: 334 mg (88%).

[**L**⁴Pd₂Cl₃] (**4a**)

To a stirred solution of **HL**⁴ (179 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), PdCl₂ (177.4 mg, 1.0 mmol) was added. This suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting orange powder was taken up in CH₂Cl₂ (40 mL) and filtered. All volatile material was then evaporated under reduced pressure to give **4a**. Yield: 276 mg (82%).

[**L**¹Ni₂Br₃]_x (**1b**)

To a stirred solution of **HL**¹ (242 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), NiBr₂(dme) (353 mg, 1.0 mmol) was added. This suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting yellow powder was taken up in 40 mL of CH₂Cl₂ and filtered. All volatile material was then evaporated under reduced pressure to give **1b**. Yield: 309 mg (74%).

[**L**²Ni₂Br₃]_x (**2b**)

To a stirred solution of **HL**² (200 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), NiBr₂(dme) (353 mg, 1.0 mmol) was added. This suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting orange powder was taken up in CH₂Cl₂ (40 mL) and filtered. All volatile material was then evaporated under reduced pressure to give **2b**. Yield: 267 mg (77%).

[**L**³Ni₂Br₃]_x (**3b**)

To a stirred solution of **HL**³ (221 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), NiBr₂(dme) (353 mg, 1.0 mmol) was added. This suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting yellow powder was taken up in CH₂Cl₂ (40 mL) and filtered. Single crystals were obtained by layering this solution with light petroleum. Yield: 223 mg (68%).

Table 3. Crystal data and refinement details for **HL¹**, **HL⁴**, **1a**, **3a**, and **3b**.

	HL¹	HL⁴	1a	3a	3b
formula	$C_{32}H_{44}N_4 \cdot 0.5 CHCl_3$	$C_{23}H_{26}N_4$	$C_{32}H_{43}Cl_3N_4Pd_2$	$C_{29}H_{37}Cl_3N_4Pd_2 \cdot 2 CH_2Cl_2 \cdot H_2O$	$C_{87}H_{111}Br_{10}N_{12}Ni_6 \cdot 6 CHCl_3$
<i>M</i>	544.40	358.48	802.85	946.63	3112.54
crystal size [mm]	$0.52 \times 0.32 \times 0.28$	$0.21 \times 0.19 \times 0.18$	$0.32 \times 0.14 \times 0.12$	$0.37 \times 0.33 \times 0.11$	$0.42 \times 0.36 \times 0.22$
crystal system	triclinic	triclinic	orthorhombic	orthorhombic	monoclinic
space group	<i>P</i> (No. 2)	<i>P</i> (No. 2)	<i>Pbca</i> (No. 61)	<i>Pccn</i> (No. 56)	<i>P2₁/n</i> (No. 14)
<i>a</i> [Å]	8.5084(1)	7.4314(13)	14.6265(3)	31.260(2)	17.3924(16)
<i>b</i> [Å]	13.2381(2)	11.8412(19)	14.7343(4)	14.4811(11)	27.051(3)
<i>c</i> [Å]	16.1319(2)	22.723(3)	32.0992(8)	16.9750(13)	26.781(3)
α [°]	112.327(1)	97.311(12)	90	90	90
β [°]	90.317(1)	90.670(13)	90	90	92.604(2)
γ [°]	106.465(1)	92.365(13)	90	90	90
<i>V</i> [Å ³]	1598.69(4)	1981.3(6)	6917.7(3)	7684.1(10)	12587(2)
$\rho_{\text{calcd.}}$ [g cm ⁻³]	1.131	1.202	1.542	1.637	1.643
<i>Z</i>	2	4	8	8	4
<i>F</i> (000)	586	768	3248	3792	6192
μ (Mo-K α) [mm ⁻¹]	0.187	0.073	1.298	1.453	4.163
$T_{\text{max}}/T_{\text{min}}$	0.9494/0.9089	0.9913/0.9841	0.8598/0.6814	1.0000/0.4246	1.0000/0.2910
<i>hkl</i> range	$\pm 10, \pm 15, -19-13$	$\pm 7, \pm 12, \pm 24$	$\pm 17, -6-16, -38-27$	$0-40, 0-19, 0-22$	$\pm 20, 0-32, 0-31$
θ range [°]	$2.52-25.00$	$1.74-22.50$	$2.06-25.00$	$2.26-27.88$	$1.39-25.00$
measured refl.	17818	7120	17074	49241	74999
unique refl. [<i>R</i> _{int}]	5534 [0.0588]	4538 [0.0878]	5633 [0.0635]	9119 [0.0538]	22161 [0.0667]
parameters	392	499	381	414	1278
goodness-of-fit	1.026	1.018	1.023	1.094	1.108
<i>R</i> ₁ / <i>wR</i> ₂ (<i>I</i> 2 σ (<i>I</i>))	0.0773/0.2027	0.1049/0.2353	0.0477/0.1063	0.0594/0.1503	0.0538/0.1396
<i>R</i> ₁ / <i>wR</i> ₂ (all data)	0.0999/0.2233	0.1747/0.2752	0.0973/0.1300	0.0916/0.1754	0.0905/0.1535
resid. el. dens. [e Å ⁻³]	0.569/−0.449	0.513/−0.404	0.605/−1.341	2.077/−1.866	2.341/−1.477

[L⁴Ni₂Br₃]_x (4b)

To a stirred solution of **HL**⁴ (179 mg, 0.5 mmol) and KO-*t*-Bu (56 mg, 0.5 mmol) in THF (50 mL), NiBr₂(dme) (353 mg, 1.0 mmol) was added. This suspension was stirred for 72 h at room temperature and then evaporated to dryness. The resulting orange powder was taken up in CH₂Cl₂ (40 mL) and filtered. All volatile material was then evaporated under reduced pressure. Yield: 287 mg (81%). According to the elemental analysis, the compound may still contain one equivalent of THF per dinickel unit.

General Procedure for Polymerisation

In a 200-mL stirred glass reactor flushed with nitrogen the appropriate amount of catalyst was added as a solution in toluene (80 mL). After the addition of MAO (for MAO/catalyst ratios see Table 1) the nitrogen atmosphere was replaced by ethylene. The pressure inside the reactor was kept at 2.5 bar for 1 hour. Finally the reaction was quenched with methanol (50 mL), and the obtained polyethylene was separated by filtration, washed with methanol and dried under vacuum to constant weight.

Polymer Characterisation

Differential scanning calorimetry was performed on a Netzsch DSC 204 F1 instrument at a heating and cooling rate of 10 K/min. DSC data reported are second heats. High temperature NMR spectra were performed in 1,1,2,2-tetrachloroethane-*d*₂ at 130 °C on a Varian Inova 400 spectrometer. Molecular weights were determined with a PL GPC-220 instrument equipped with mixed B columns in trichlorobenzene at 160 °C vs. linear polyethylene standards.

X-Ray Crystallographic Studies

X-ray data were collected on a four circle CCD (**1a**, **HL**⁴), a STOE IPDS II (**HL**⁴), and a Bruker AXS CCD diffractometer (**3a**, **3b**) (graphite monochromated Mo-K α radiation, λ = 0.71073 Å) at –140 °C (**1a**, **HL**⁴, **HL**⁴) and –83 °C (**3a**, **3b**) (Table 3). The structures were solved by direct methods and refined on F^2 using all reflections with SHELX-97.^[25] The non-hydrogen atoms were refined anisotropically. Hydrogen atoms which were not involved in hydrogen bonding were placed in calculated positions and assigned to an isotropic displacement parameter of 0.08 Å² (**1a**, **HL**⁴, **HL**⁴), or included using the riding model (**3a**, **3b**). The positional and isotropic thermal parameters of the nitrogen-bonded hydrogen atom H(2) in **HL**⁴ and **HL**⁴ were refined without constraints. SADABS was used to perform area-detector scaling and absorption corrections for **1a**, **3a**, **3b**, and **HL**⁴.^[26] Face-indexed absorption corrections were performed numerically with the program X-RED for **HL**⁴.^[27] Disordered solvent molecules were found in **3b** (CHCl₃) and were refined in the usual manner, as well as the disordered solvent molecules on special positions in **1a** (CHCl₃) and **3a** (CH₂Cl₂, H₂O). For some solvent molecules in **3b** no satisfactory model for the disorder could be found, and for further refinement the contribution of the missing solvent molecules (980 Å³, total electron count 210) was subtract-

ed from the reflection data by the SQUEEZE^[28] routine of the PLATON^[29] program. Additionally the methyl groups [C(28), C(29)] of one isopropyl side arm in **HL**⁴ are disordered about two positions. The respective occupancy factors were refined to 0.52(3)/0.48(3).

CCDC-256780 (**HL**⁴), CCDC-256781 (**HL**⁴), CCDC-256782 (**1a**), CCDC-256783 (**3a**), and CCDC-256784 (**3b**) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge at www.ccdc.cam.ac.uk/conts/retrieving.html [or from the Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB2 1EZ, UK; fax: (internat.) +44-1223/336-033; E-mail: deposit@ccdc.cam.ac.uk].

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